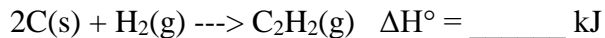


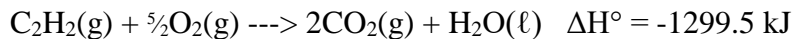
Enthalpy Questions (Using Hess' Law)

Name \_\_\_\_\_

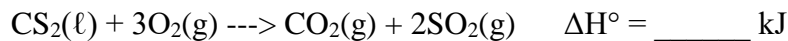
**#1:** Calculate the enthalpy for this reaction:



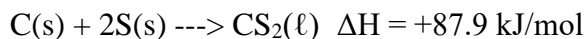
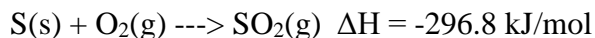
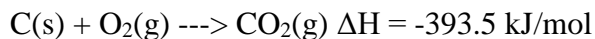
Given the following thermochemical equations:



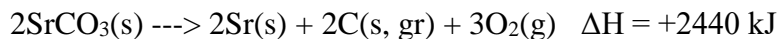
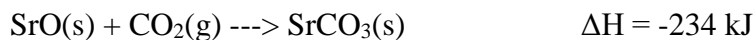
**#2:** Calculate the enthalpy of the following chemical reaction:



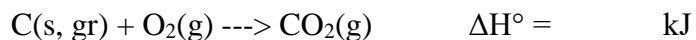
Given:



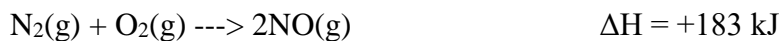
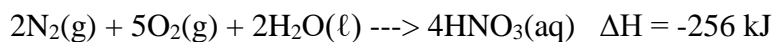
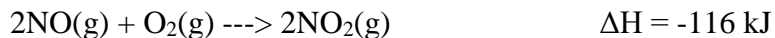
**#3:** Given the following data:



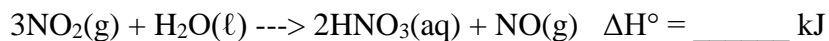
Find the  $\Delta H$  of the following reaction:



**#4:** Given the following information:

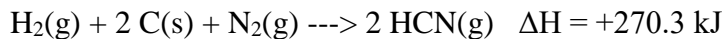
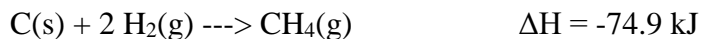
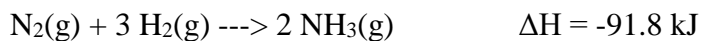


Calculate the enthalpy change for the reaction below:

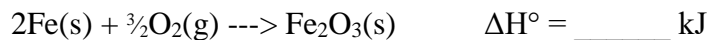


**#5:** Calculate  $\Delta H$  for this reaction:  $\text{CH}_4(\text{g}) + \text{NH}_3(\text{g}) \rightarrow \text{HCN}(\text{g}) + 3\text{H}_2(\text{g}) \quad \Delta H^\circ = \underline{\hspace{2cm}} \text{ kJ}$

Given:



**#6:** Determine the heat of reaction for the oxidation of iron:



Given the thermochemical equations:

